

## Electron Configurations - Solutions

*Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different.*

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
- 6) cobalt  $[\text{Ar}] 4s^2 3d^7$
- 7) silver  $[\text{Kr}] 5s^2 4d^9$
- 8) tellurium  $[\text{Kr}] 5s^2 4d^{10} 5p^4$
- 9) radium  $[\text{Rn}] 7s^2$
- 10) lawrencium  $[\text{Rn}] 7s^2 5f^{14} 6d^1$
- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  rubidium
- 13)  $[\text{Kr}] 5s^2 4d^{10} 5p^3$  antimony
- 14)  $[\text{Xe}] 6s^2 4f^{14} 5d^6$  osmium
- 15)  $[\text{Rn}] 7s^2 5f^{11}$  einsteinium
- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (take a look at "4d")
- 17)  $1s^2 2s^2 2p^6 3s^3 3d^5$  not valid (3p comes after 3s)
- 18)  $[\text{Ra}] 7s^2 5f^8$  not valid (radium isn't a noble gas)
- 19)  $[\text{Kr}] 5s^2 4d^{10} 5p^5$  valid
- 20)  $[\text{Xe}]$  not valid (an element can't be its own electron configuration)

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- 1) oxygen  $1s^2 2s^2 2p^4$
- 2) sodium  $1s^2 2s^2 2p^6 3s^1$
- 3) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 4) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 5) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 6) nitrogen  $1s^2 2s^2 2p^3$
- 7) chlorine  $1s^2 2s^2 2p^6 3s^2 3p^5$
- 8) argon  $1s^2 2s^2 2p^6 3s^2 3p^6$
- 9) cobalt  $[\text{Ar}] 4s^2 3d^7$
- 10) silver  $[\text{Kr}] 5s^2 4d^9$
- 11) tellurium  $[\text{Kr}] 5s^2 4d^{10} 5p^4$
- 12) iodine  $[\text{Kr}] 5s^2 4d^{10} 5p^5$
- 13) cesium  $[\text{Xe}] 6s^1$
- 14)  $1s^2 2s^2 2p^6 3s^2 3p^4$  sulfur
- 15)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  rubidium
- 16)  $[\text{Kr}] 5s^2 4d^{10} 5p^3$  antimony
- 17)  $[\text{Xe}] 6s^2 4f^{14} 5d^6$  osmium
- 18)  $[\text{Xe}] 6s^2$  barium

These electron configurations have mistakes, determine what is wrong.

- 19)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (take a look at "4d")
- 20)  $1s^2 2s^2 2p^6 3s^2 3d^5$  not valid (3p comes after 3s)
- 21)  $[\text{Ra}] 7s^2 5f^8$  not valid (radium isn't a noble gas)
- 22)  $[\text{Xe}]$  not valid (an element can't be its own electron configuration)